**1**

The word equation below shows a reaction used in an industrial process.

chromium oxide + aluminium chromium + aluminium oxide

 The reaction is highly exothermic.

(a)What is an exothermic reaction?

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**(2)**

(b)Name the products of this reaction.

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**(1)**

(c)What does the term Endothermic mean?

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.................................................................................................................... (2)

(d) When potassium nitrate dissolves in water, the beaker containing the solution gets cooler. Is dissolving this salt an exothermic or an endothermic process?

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.................................................................................................................... (2)

 (e) Name a reaction that is Exothermic

 ………………………………………………………………………….. (1)

 (f) Name a reaction that is Endothermic

 ……………………………………………………………………………. (1)

 (g) An acid was added to an alkali. The temperature at the start was 21 degrees celcius and afterwards the temperature was 30 degrees celcius. Select 2 words to describe the reaction-

Exothermic Endothermic Combustion Respiration Neutralisation Acidification (2)

**Total 11 marks**